Ссылка: <http://85.142.162.117/os/xmodules/qprint/openlogin.php?fvq=4EB05C51C08F9EAF478482D229AC845B>

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  H2SO4+ Al → Al2(SO4)3 + S + H2O  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | F3DB40 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  Ag + HClO3 → AgCl + AgClO3 + H2O  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 04114B | |
| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  NaMnO4 + NaOH → Na2MnO4 + O2 + H2O  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 56E240 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  HBrO3 + H2S → S + Br2 + H2O  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | AFB445 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  Zn + HNO3 → Zn(NO3)2 + N2O + H2O  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | A48C46 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  Br2 + KI + H2O → KIO3 + HBr  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | E4BF4D | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  S + KOH → K2S + K2SO3 + H2O  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | E8AF45 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  H2SO4 + Al →H2S + Al2(SO4)3 + H2O  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 88484E | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  K2S + HNO3 → K2SO4 + NO + H2O  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 05D5F3 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  K2S + HClO3 + H2O→ KCl + S + KOH  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | B487FC | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  Ca + HNO3 → N2O + Ca(NO3)2 + H2O  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | D8D5F7 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  PH3 + Cl2 + H2O → H3PO4 + HCl  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 5C29FC | |
| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  H2SO4+ Fe → Fe2(SO4)3 + SO2 + H2O   Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 90E9FA | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  MnCO3 + KClO3 → MnO2 + KCl + CO2  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 913CF9 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  HNO3 + Na2SO3 → Na2SO4 + NO2 + H2O  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 97DDF3 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  Cr + O2 + HCl → CrCl3 + H2O  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | E142F9 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  B + HBrO3 + H2O → H3BO3 + HBr  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | E48DF3 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  CrO3 + NH3 → Cr2O3 + N2 + H2O  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | E250F8 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  HNO2 + HI → I2 + NO + H2O  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 694DFB | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  HNO3 + PbS → PbSO4 + NO + H2O  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 37F5F5 | |

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| |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | Начало формы   |  | | --- | | Взаимодействию карбоната магния с бромоводородной кислотой соответствует сокращённое ионное уравнение | | |  |  |  | | --- | --- | --- | |  | **1)** | MgCO3 + 2H+ = Mg2+ + CO2 + H2O | |  | **2)** | CO32– + 2H+ = CO2 + H2O | |  | **3)** | MgCO3 + 2H+ + 2Br– = MgBr2 + H2O + CO2 | |  | **4)** | MgCO3 + 2HBr = Mg2+ + 2Br– + H2O + CO2 | |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 10948A | |

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| |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | Начало формы   |  | | --- | | Реакции соединения, протекающей без изменения степеней окисления, соответствует схема | | |  |  |  | | --- | --- | --- | |  | **1)** | NH3 + HNO3 → NH4NO3 | |  | **2)** | S + Na2SO3 → Na2S2O3 | |  | **3)** | CO + Na2O2 → Na2CO3 | |  | **4)** | PH3 + O2 → H3PO4 | |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | B9E519 | |

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| |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | Начало формы   |  | | --- | | Осадок **не образуется** при взаимодействии растворов | | |  |  |  | | --- | --- | --- | |  | **1)** | гидроксида бария и фосфата натрия | |  | **2)** | гидроксида натрия и фосфорной кислоты | |  | **3)** | нитрата серебра и иодида калия | |  | **4)** | фосфата калия и нитрата цинка | |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 4D7E60 | |

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| |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | | Начало формы   |  | | --- | | Сокращённому ионному уравнению 2H+ + CO32– = CO2 + H2O соответствует левая часть схемы реакции | | |  |  |  | | --- | --- | --- | |  | **1)** | H2SiO3 + K2CO3 → | |  | **2)** | H2O + CO2 → | |  | **3)** | HCl + Na2CO3 → | |  | **4)** | HNO3 + CaCO3 → | |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 499160 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  Fe(OH)3 + HI → FeI2 + I2 + H2O  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 7AA907 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  AgNO3 + Cl2 → AgCl + O2 + N2O5  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 7CE906 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  HClO4 + SO2 + H2O → HCl + H2SO4  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | DA3605 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  K2MnO4 + H2O → MnO2 + KMnO4 +KOH  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | A34203 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  HCl + PbO2 → PbCl2 + Cl2 + H2O  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | CFFA00 | |
| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  H2O2 + HIO3 → O2 + I2 + H2O  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | FAEB7D | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  Na2S + FeCl3 → FeS + S + NaCl  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 75A879 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  HBr + KBrO3 → Br2 + KBr + H2O  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | BB7373 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  Na2S + FeBr3 → FeS + S + NaBr  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | BF3573 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  HNO2+ Cl2 + H2O → HNO3 + HCl  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | D87D73 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  S + F2 + NaOH → Na2SO4 + NaF + H2O  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | AF857F | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  H2O2 + KClO3 → KCl + O2 + H2O  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 8CFB77 | |

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| |  |  | | --- | --- | | Начало формы   |  | | --- | | Используя метод электронного баланса, составьте уравнение реакции  HBr + MnO2 → MnBr2 + Br2 + H2O  Определите окислитель и восстановитель. |   Конец формы |  |  |  | | --- | --- | | http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 86907C | |