Ссылка: <http://85.142.162.117/os/xmodules/qprint/openlogin.php?fvq=4EB05C51C08F9EAF478482D229AC845B>

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| Начало формы

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| Уравнению окислительно-восстановительной экзотермической реакции соответствует запись |
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|    |  **1)**  | CO2 + Na2O2 → Na2CO3 + O2 + Q |
|    |  **2)**  | Al(OН)3 → Al2O3 + H2O – Q |
|    |  **3)**  | 2NH3 → N2 + 3H2– Q |
|    |  **4)**  | Na2O + P2O5 → Na3PO4+ Q |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | FB9874 |

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| Начало формы

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| Реакции соединения соответствует уравнение |
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|    |  **1)**  | FeO + H2 = Fe + H2O |
|    |  **2)**  | H2O + Cl2 = HClO + HCl |
|    |  **3)**  | P2O5 + 3H2O = 2H3PO4 |
|    |  **4)**  | CaO + 2HCl = CaCl2 + H2O |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 419E1F |

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| Начало формы

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| Установите соответствие между формулами двух веществ и реактивом, с помощью которого можно различить эти вещества. |

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| **ФОРМУЛЫ ВЕЩЕСТВ** |   | **РЕАКТИВ** |
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| **А)**  | HCl и HNO3 |
| **Б)**  | Bа(NO3)2 и Na2SO4 |
| **В)**  | KCl и NH4Cl |

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| **1)**  | соляная кислота |
| **2)**  | карбонат калия |
| **3)**  | медь |
| **4)**  | гидроксид натрия |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | A90B9D |

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| Начало формы

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| Установите соответствие между формулами двух веществ и реактивом, с помощью которого можно различить эти вещества. |

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| **ФОРМУЛЫ ВЕЩЕСТВ** |   | **РЕАКТИВ** |
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| --- | --- |
| **А)**  | KCl и NaNO3 |
| **Б)**  | Вa(NO3)2 и KNO3 |
| **В)**  | NH4Cl и AlCl3 |

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| **1)**  | гидроксид натрия |
| **2)**  | сульфат натрия |
| **3)**  | оксид магния |
| **4)**  | нитрат серебра |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | E98E98 |

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| Начало формы

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| Взаимодействию гидроксида меди(II) с раствором серной кислоты соответствует сокращённое ионное уравнение |
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|    |  **1)**  | Cu2+ + 2OH– = Cu(OH)2 |
|    |  **2)**  | Сu2+ + SO42– = СuSO4 |
|    |  **3)**  | Cu(OH)2 + 2H+ = Cu2+ + 2H2O |
|    |  **4)**  | Cu(OH)2 + SO42– = CuSO4 + 2OH– |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 1ED467 |

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| Начало формы

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| Сокращённое ионное уравнение Fe2+ + 2OH− = Fe(OH)2 соответствует взаимодействию между  |
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|    |  **1)**  | нитратом железа(II) и гидроксидом калия |
|    |  **2)**  | сульфатом железа(II) и гидроксидом бария |
|    |  **3)**  | фосфатом железа(II) и водой |
|    |  **4)**  | оксидом железа(II) и водой |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 004C41 |

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| Начало формы

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| Реакции обмена соответствует левая часть схемы реакции |
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|    |  **1)**  | NaI + Pb(NO3)2 → |
|    |  **2)**  | Cu + O2 → |
|    |  **3)**  | С + H2SO4(конц) → |
|    |  **4)**  | Fe2O3 + CO → |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 2211C5 |

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| Начало формы

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| Сокращённое ионное уравнение3Ag+ + PO43– = Ag3PO4соответствует взаимодействию |
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|    |  **1)**  | Ag2O и K3PO4 |
|    |  **2)**  | Ag и H3PO4 |
|    |  **3)**  | AgNO3 и K3PO4 |
|    |  **4)**  | Ag2O и H3PO4 |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | F1E077 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииH2SO4+ Al → Al2(SO4)3 + S + H2OОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | F3DB40 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииAg + HClO3 → AgCl + AgClO3 + H2OОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 04114B |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииNaMnO4 + NaOH → Na2MnO4 + O2 + H2OОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 56E240 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииHBrO3 + H2S → S + Br2 + H2OОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | AFB445 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииZn + HNO3 → Zn(NO3)2 + N2O + H2OОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | A48C46 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииBr2 + KI + H2O → KIO3 + HBrОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | E4BF4D |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииS + KOH → K2S + K2SO3 + H2OОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | E8AF45 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииH2SO4 + Al →H2S + Al2(SO4)3 + H2OОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 88484E |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииK2S + HNO3 → K2SO4 + NO + H2OОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 05D5F3 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииK2S + HClO3 + H2O→ KCl + S + KOHОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | B487FC |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииCa + HNO3 → N2O + Ca(NO3)2 + H2OОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | D8D5F7 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииPH3 + Cl2 + H2O → H3PO4 + HClОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 5C29FC |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииH2SO4+ Fe → Fe2(SO4)3 + SO2 + H2O Определите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 90E9FA |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииMnCO3 + KClO3 → MnO2 + KCl + CO2Определите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 913CF9 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииHNO3 + Na2SO3 → Na2SO4 + NO2 + H2OОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 97DDF3 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииCr + O2 + HCl → CrCl3 + H2OОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | E142F9 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииB + HBrO3 + H2O → H3BO3 + HBrОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | E48DF3 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииCrO3 + NH3 → Cr2O3 + N2 + H2OОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | E250F8 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииHNO2 + HI → I2 + NO + H2OОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 694DFB |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииHNO3 + PbS → PbSO4 + NO + H2OОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 37F5F5 |

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| Начало формы

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| Сокращённое и полное ионное уравнения совпадают для реакции между |
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|    |  **1)**  | AgNO3 и KCl |
|    |  **2)**  | CuSO4 и NaOH |
|    |  **3)**  | HCl и CaCO3 |
|    |  **4)**  | H2SO4 и Ba(OH)2 |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | C8052B |

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| Начало формы

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| Реакция нейтрализации происходит при взаимодействии веществ |
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| --- | --- | --- |
|    |  **1)**  | BaO и H2SO4 |
|    |  **2)**  | KOH и HNO3 |
|    |  **3)**  | Na2O и H2O |
|    |  **4)**  | CaCO3 и HCl |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | ECF536 |

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| Начало формы

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| Взаимодействию карбоната магния с бромоводородной кислотой соответствует сокращённое ионное уравнение |
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|    |  **1)**  | MgCO3 + 2H+ = Mg2+ + CO2 + H2O  |
|    |  **2)**  | CO32– + 2H+ = CO2 + H2O  |
|    |  **3)**  | MgCO3 + 2H+ + 2Br– = MgBr2 + H2O + CO2  |
|    |  **4)**  | MgCO3 + 2HBr = Mg2+ + 2Br– + H2O + CO2  |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 10948A |

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| Начало формы

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| --- |
| Сокращённое ионное уравнение 2H+ + FeS = H2S + Fe2+ соответствует взаимодействию сульфида железа и |
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|    |  **1)**  | фосфорной кислоты |
|    |  **2)**  | соляной кислоты |
|    |  **3)**  | кремниевой кислоты |
|    |  **4)**  | сернистой кислоты |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | F8E18E |

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| Начало формы

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| --- |
| Газ выделяется при взаимодействии |
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| --- | --- | --- |
|    |  **1)**  | MgCl2 и Ba(NO3)2 |
|    |  **2)**  | Na2CO3 и CaCl2 |
|    |  **3)**  | NH4Cl и NaOH |
|    |  **4)**  | CuSO4 и KOH |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 2453CE |

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| Начало формы

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| --- |
| Осадок образуется при взаимодействии водных растворов |
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| --- | --- | --- |
|    |  **1)**  | нитрата серебра и бромида кальция |
|    |  **2)**  | нитрата аммония и гидроксида калия |
|    |  **3)**  | соляной кислоты и гидроксида бария |
|    |  **4)**  | азотной кислоты и карбоната натрия |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 1D3C9F |

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| Начало формы

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| Окислительно-восстановительной реакции соединения соответствует схема реакции |
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| --- | --- | --- |
|    |  **1)**  | SO2 + H2O → H2SO3 |
|    |  **2)**  | H2S + Ca(OH)2 → CaS + H2O |
|    |  **3)**  | H2SO4 + Zn → ZnSO4 + H2 |
|    |  **4)**  | SO2 + H2O2 → H2SO4 |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 03CFE8 |

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| Начало формы

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| --- |
| Осадок образуется при взаимодействии водных растворов |
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|    |  **1)**  | NaNO3 и CaCl2 |
|    |  **2)**  | Na2S и AgNO3 |
|    |  **3)**  | MgBr2 и HCl |
|    |  **4)**  | (NH4)3PO4 и KOH |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | E15629 |

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| Начало формы

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| Реакции соединения, протекающей без изменения степеней окисления, соответствует схема реакции |
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| --- | --- | --- |
|    |  **1)**  | P2O3 + O2 → P2O5 |
|    |  **2)**  | NaOH + SO2 → NaHSO3 |
|    |  **3)**  | NO2 + O2 + H2O → HNO3 |
|    |  **4)**  | NH3 + O2 → NO + H2O |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | A53F96 |

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| Начало формы

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| Газ **не выделяется** в ходе реакции между растворами |
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|    |  **1)**  | HNO3 и K2CO3 |
|    |  **2)**  | (NH4)2SO4 и Ba(OH)2 |
|    |  **3)**  | Na2CO3 и Ba(OH)2 |
|    |  **4)**  | HCl и Na2S |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 61B531 |

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| Начало формы

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| Окислительно-восстановительной реакции соединения соответствует схема |
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| --- | --- | --- |
|    |  **1)**  | BaO + H2O → Ba(OH)2 |
|    |  **2)**  | P2O5 + CaO → Ca3(PO4)2 |
|    |  **3)**  | NO2 + H2O + O2 → HNO3 |
|    |  **4)**  | Cl2O7 + Na2O → NaClO4 |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 1AA06D |

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| Начало формы

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| Осадок **не образуется** в ходе реакции между растворами |
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|    |  **1)**  | FeSO4 и KOH |
|    |  **2)**  | Na3PO4 и AgNO3 |
|    |  **3)**  | H2SO4 и Ba(OH)2 |
|    |  **4)**  | H2SO4 и NaOH |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | F4B9A2 |

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| Начало формы

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| Реакции соединения, протекающей без изменения степеней окисления, соответствует схема |
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|    |  **1)**  | NH3 + HNO3 → NH4NO3 |
|    |  **2)**  | S + Na2SO3 → Na2S2O3 |
|    |  **3)**  | CO + Na2O2 → Na2CO3 |
|    |  **4)**  | PH3 + O2 → H3PO4 |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | B9E519 |

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| Начало формы

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| Реакция ионного обмена идёт практически до конца между растворами |
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|    |  **1)**  | Ca(NO3)2 и K3PO4 |
|    |  **2)**  | KCl и H2SO4 |
|    |  **3)**  | NaNO3 и MgCl2 |
|    |  **4)**  | K2SO4 и NaOH |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 985496 |

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| Начало формы

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| В реакцию обмена вступают вещества |
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|    |  **1)**  | Na2O и SO3 |
|    |  **2)**  | CaO и H2O |
|    |  **3)**  | K2S и HCl |
|    |  **4)**  | Zn и H2SO4 |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 92DD1A |

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| Начало формы

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| Практически до конца протекает реакция ионного обмена между |
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|    |  **1)**  | гидроксидом натрия и нитратом бария |
|    |  **2)**  | гидроксидом калия и сульфатом цинка |
|    |  **3)**  | фосфатом калия и сульфатом натрия |
|    |  **4)**  | нитратом калия и соляной кислотой |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 1B7641 |

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| Начало формы

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| --- |
| Осадок образуется при взаимодействии ионов |
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| --- | --- | --- |
|    |  **1)**  | H+ и NO3–  |
|    |  **2)**  | Ba2+ и OH–  |
|    |  **3)**  | Ca2+ и PO43–  |
|    |  **4)**  | Al3+ и SO42–  |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 6A0FAF |

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| Начало формы

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| --- |
| В реакцию соединения вступают вещества |
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| --- | --- | --- |
|    |  **1)**  | BaO и HCl |
|    |  **2)**  | K2O и P2O5 |
|    |  **3)**  | CaO и H2S |
|    |  **4)**  | NH3 и O2 |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 8B9232 |

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| Начало формы

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| В водном растворе одновременно могут находиться ионы |
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|    |  **1)**  | H+ и CO32– |
|    |  **2)**  | Cu2+ и OH– |
|    |  **3)**  | Mg2+ и OH– |
|    |  **4)**  | Na+ и CO32– |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | C6BC9E |

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| Начало формы

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| Признаком протекания химической реакции между оксидом меди и водородом является |
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| --- | --- | --- |
|    |  **1)**  | появление запаха |
|    |  **2)**  | изменение цвета |
|    |  **3)**  | выпадение осадка |
|    |  **4)**  | выделение газа |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 2CC5D3 |

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| Начало формы

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| Газ выделяется при взаимодействии водных растворов |
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|    |  **1)**  | фосфата калия и соляной кислоты |
|    |  **2)**  | сульфата натрия и хлорида бария |
|    |  **3)**  | гидроксида калия и серной кислоты |
|    |  **4)**  | карбоната натрия и соляной кислоты |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | E67862 |

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| Начало формы

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| Признаком протекания химической реакции между гидроксидом меди(II) и соляной кислотой является |
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|    |  **1)**  | растворение осадка |
|    |  **2)**  | выпадение осадка |
|    |  **3)**  | появление запаха |
|    |  **4)**  | выделение газа |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 7AA271 |

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| Начало формы

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| Осадок **не образуется** при взаимодействии растворов |
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|    |  **1)**  | гидроксида бария и фосфата натрия |
|    |  **2)**  | гидроксида натрия и фосфорной кислоты |
|    |  **3)**  | нитрата серебра и иодида калия |
|    |  **4)**  | фосфата калия и нитрата цинка |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 4D7E60 |

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| Начало формы

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| Сумма коэффициентов в уравнении реакции между карбонатом кальция и соляной кислотой равна |
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|    |  **1)**  | 6 |
|    |  **2)**  | 7 |
|    |  **3)**  | 8 |
|    |  **4)**  | 9 |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | DDA588 |

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| Начало формы

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| Сокращённому ионному уравнению 2H+ + CO32– = CO2 + H2O соответствует левая часть схемы реакции |
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|    |  **1)**  | H2SiO3 + K2CO3 → |
|    |  **2)**  | H2O + CO2 → |
|    |  **3)**  | HCl + Na2CO3 → |
|    |  **4)**  | HNO3 + CaCO3 → |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 499160 |

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| Начало формы

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| Сумма коэффициентов в уравнении реакции между натрием и водой равна |
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|    |  **1)**  | 6 |
|    |  **2)**  | 7 |
|    |  **3)**  | 8 |
|    |  **4)**  | 10 |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 6EDF7B |

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| Начало формы

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| Реакция ионного обмена протекает практически до конца при взаимодействии |
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|    |  **1)**  | хлороводорода и сульфата алюминия |
|    |  **2)**  | хлороводорода и нитрата кальция |
|    |  **3)**  | нитрата кальция и хлорида натрия |
|    |  **4)**  | нитрата серебра и хлороводорода |

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Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | BD6EA1 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииFe(OH)3 + HI → FeI2 + I2 + H2OОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 7AA907 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииAgNO3 + Cl2 → AgCl + O2 + N2O5Определите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 7CE906 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииHClO4 + SO2 + H2O → HCl + H2SO4Определите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | DA3605 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииK2MnO4 + H2O → MnO2 + KMnO4 +KOHОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | A34203 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииHCl + PbO2 → PbCl2 + Cl2 + H2OОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | CFFA00 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииH2O2 + HIO3 → O2 + I2 + H2OОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | FAEB7D |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииNa2S + FeCl3 → FeS + S + NaClОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 75A879 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииHBr + KBrO3 → Br2 + KBr + H2OОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | BB7373 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииNa2S + FeBr3 → FeS + S + NaBrОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | BF3573 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииHNO2+ Cl2 + H2O → HNO3 + HClОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | D87D73 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииS + F2 + NaOH → Na2SO4 + NaF + H2OОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | AF857F |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииH2O2 + KClO3 → KCl + O2 + H2OОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 8CFB77 |

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| Начало формы

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| Используя метод электронного баланса, составьте уравнение реакцииHBr + MnO2 → MnBr2 + Br2 + H2OОпределите окислитель и восстановитель. |

Конец формы |

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| http://85.142.162.117/os/images/delfav.gif http://85.142.162.117/os/admin/images/view0.gif | 86907C |

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